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Chemistry Chapter 12

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Remember it is a MC test,

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2 of 2) 19. The mass of

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element X found in 1.00 mole of each of four different compounds is 28.0 g, 42.0 g, 56.0 g, and 70 g, respectively. The possible atomic weight of X is a. 8.00 b. 14.0

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Answers: Moles and
Stoichiometry Practice
Problems 1) How many moles
of sodium atoms correspond
to 1.56×10^{21} atoms of
sodium? 1.56×10^{21} atoms
Na \times $1 \text{ mol Na} = 2.59 \times 10^{-3}$
 mol Na $236.022 \times 10 \text{ atoms Na}$
2) Determine the mass in
grams of each of the
following: a. 1.35 mol of Fe
 $1.35 \text{ mol Fe} \times 55.845 \text{ g Fe} =$
 75.4 g Fe 1 mol Fe b. 24.5
 mol O

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Chapter 9 Stoichiometry Test Answers

SOLUTION: You are asked to calculate the amount (in moles) of oxygen atoms in a

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given sample. You are
given the mass of the oxygen
sample. Use the molar mass
of oxygen (1 mol O = 16.00
g) to create a conversion
factor that converts mass
(in grams) to amount (in
mol) of oxygen. $124 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g}} = 7.75 \text{ mol O}$.

Chapter 3 Stoichiometry

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(atoms, (molecules) ((2.
1 mole (= 6.022×10^{23} particles
((1 mole = molar (mass
(1 mole = 22.4 L (3.

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