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surfaces affect speed.

o Except for substances in the gaseous state or solution, reactions occur at the boundary, or interface, between two phases.

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a 61-slide Power

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Point presentation
and 11-page guided
notes packet used to
explain some of the
fundamental concepts
of chemical kinetics.
It covers various
reaction rate topics
including: zero, first,
and second order
rates, calculating
using the method of
initial rates, half lives,
colli.

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TpT Worksheet

$2\text{O}_2(\text{g}) \rightarrow 2\text{N}_2\text{O}(\text{g}) + \text{O}_2(\text{g})$

rate = $k[\text{N}_2\text{O}]^2$

For an initial

concentration of N_2O

of 0.50 M,

calculate the

concentration of N_2O

remaining after

2.0 min if $k = 3.4$

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x10⁻³ s⁻¹. 15. A

reaction has a rate constant of 0.00300 s⁻¹.

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Chemistry Period:
Date: Dr. Mandes The
following questions
represent potential
types of quiz

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questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday. 5. Please do #18 in chapter 12 of your text. a.

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Problems and
Solutions

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CHOICE. Choose the

one alternative that

best completes the

statement or answers

the question. 1)

Consider the

following reaction: $3A$

$\rightarrow 2B$ The average rate

of appearance of B is

given by $D[B]/Dt$.

Comparing the rate of

appearance of B and

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the rate of disappearance of A, we get $D[B]/Dt =$
 $\times (-D[A]/Dt)$.

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The rate equation for a chemical reaction is determined by (A)

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theoretical

calculations. (B)

measuring reaction

rate as a function of

concentration of

reacting species. (C)

determining the

equilibrium constant

for the reaction. (D)

measuring reaction

rates as a function of

temperature 2.

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Response Questions

page 2 Mechanism 3

is correct. The rate

law shows that the

slow reaction must

involve one Y,

consistent with

mechanism 3.

Mechanisms 1 and 2

would involve both

[X] and [Y] in the rate

law, not consistent

with the rate law.

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1987 a) three points;
one each for form of
rate law, HgCl_2
exponent, $\text{C}_2\text{O}_4^{2-}$
exponent

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In this kinetics
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students find rate laws of reactions, they determine rate constants, they find rate-determining steps of reactions, they find half-life's of reactions and they determine the overall order of reactions.

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are drawn into a
vacuum where the
following process
takes place. F
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– the study of rates
of chemical reactions
and the mechanisms
by which they occur

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react faster (ie ...

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The AP Chemistry

course is taught as a

continuation of the

first year course. As

such, units are

numbered

accordingly. Units 1-8

are covered during

the first year course

and reviewed over a

short time span

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during the AP course.
(Links for these units
can be found under
the "CHEMISTRY" tab
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6 Kinetics and
Equilibrium Chemical
kinetics can be

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divided into two parts. The first, at the macroscopic level, is the study of rates of reactions: what the rate of reaction means; how to determine a rate by experiment; and how factors, such as the concentrations of reactants and temperature, influence rates.

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